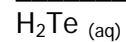
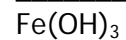
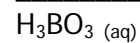
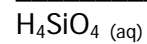
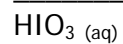
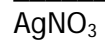
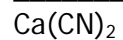
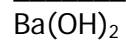


NAMING COMPOUNDS

PRACTICE



NAMING COMPOUNDS

PRACTICE ~ ANSWER KEY

SO ₂	Sulfur dioxide
CuCl	Copper (I) chloride
Ba(OH) ₂	Barium hydroxide
FeS	Iron (II) sulfide
N ₂ O ₄	Dinitrogen tetroxide
As ₄ O ₆	Tetraarsenic hexoxide
Ca(CN) ₂	Calcium cyanide
NH ₄ I	Ammonium iodide
SnCl ₄	Tin (IV) chloride
CaCl ₂	Calcium chloride
AgNO ₃	Silver nitrate
NaH	Sodium hydride
Cl ₂ O ₇	Dichlorine heptoxide
Al ₂ O ₃	Aluminum oxide
Rb ₂ S	Rubidium sulfide
LiF	Lithium Fluoride
FeBr ₃	Iron (III) bromide
HIO ₃ (aq)	Iodic acid

FeO	Iron (II) oxide
PbS ₂	Lead (IV) sulfide
Bi ₂ O ₅	Bismuth (V) oxide
KBr	Potassium Bromide
CuCl ₂	Copper (II) chloride
SrO	Strontium oxide
H ₄ SiO ₄ (aq)	Silicic acid
HCl (aq)	Hydrochloric acid
H ₃ BO ₃ (aq)	Bromic acid
HCN	Hydrogen cyanide
H ₂ S (aq)	Hydrosulfuric acid
Fe(OH) ₃	Iron (III) hydroxide
CS ₂	Carbon disulfide
SO ₃	Sulfur trioxide
Cr ₂ O ₃	Chromium (III) oxide
H ₂ Te (aq)	Hydrotelluric acid
Ba ₃ N ₂	Barium nitride
CuF ₂	Copper (II) fluoride